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Masters Chemistry: Matter and Change □

Chapter 11 31 number of moles of known
substance C H C H C H O O O CO CO

CO H O H O H O number of moles of

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Chapter 13.4
unknown substance O CO H O C H CO H
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Chemistry: Matter and Change □ Chapter 4

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Calculating Atomic Mass MATH SKILLS

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the table, calculate the atomic mass of

unknown element X. Then, identify the

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4.3. 4. Given the data in the table,
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element X. Then, identify the unknown element, which is used medically to treat some mental disorders. Analyze the Problem. You are given the data in the table.

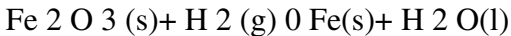
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Chapter 11, 1.2 MATH SKILLS

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Explain what a crystalline solid is. 18 Use with Chapter 12, Section 2.3 1. The reaction of iron(III) oxide (Fe_2O_3) and hydrogen (H_2) is represented by the following unbalanced chemical equation.



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Chapter 2: Analyzing Data

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WORKSHEET Determining Numbers of
Ions Use with Chapter 7, Section 7.3 1.

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Write a simple word equation that illustrates what must be true of total positive charge and total negative charge in an ionic compound. 2. How many potassium ions (group 1) would be related to balance the charge of each of the following in a compound?

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June 13th, 2018 - 94 Chemistry Matter and
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82 Chemistry: Matter and Change □
Chapter 25 Math Skills Transparency
Worksheets 1. In equation 1, what do the

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numbers 238 and 92, written to the left of the symbol U, represent? 2. Explain how both mass number and atomic number have been conserved in equation 1. 3. Write the nuclear symbol for the missing term in equation 2. 4.

balancing nuclear equations

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Chapter 13 1 Math Skills Transparency Worksheets

Chemistry: Matter and Change □ Chapter 5

13 1. Look at the two waves shown. What is the speed of each wave? 2. Look at the two waves shown. Which wave has a higher frequency? Which wave has a longer wavelength? 3. Assume that wave A has a wavelength of 699 nm. Calculate

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the frequency of the wave. Show your work. 4.

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WORKSHEET** Use with Chapter 25,

Section 25.3 42 82 1. In equation 1, what
do the numbers 238 and 92, written to the
left of the symbol \cup , represent? 238 is the

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Chapter 13.1
mass number of the isotope, and 92 is the atomic number of the isotope. 2. Explain how both mass number and atomic number have been conserved in equation 1.

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Worksheet Answers MATH SKILLS

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Chapter 7, Section 10.1 73 1 Write a simple word equation that illustrates what must be true of total positive charge and total negative charge in an ionic compound 2 How many potassium ions (group 1)

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