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The initial  $[H_2]$  is 0.0420mol/L and the  
initial  $[CO_2]$  is 0.0160mol/L. At

equilibrium, the  $[CO_2]$  is 0.0139mol/L.

Calculate the equilibrium constant for the

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several factors: temperature, concentration, nature of the reactants, surface area and the presence of a catalyst.

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statement or answers the question. 1)

Consider the following reaction:  $3A \rightarrow 2B$   
The average rate of appearance of B is

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Answers given by  $D[B]/Dt$ . Comparing the rate of appearance of B and the rate of disappearance of A, we get  $D[B]/Dt =$  \_\_\_\_\_ x  $(-D[A]/Dt)$ .

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kinetics and chemical equilibrium. Adult learners enrolled in this course study phenomena or technological applications related to the reaction rate of a chemical reaction or to chemical equilibrium, and look for answers or solutions to problems involving them. They thus acquire knowledge about rate law, equilibrium constants, factors

### **Chemistry: Kinetics and Equilibrium**

Regents review Kinetics & equilibrium A) decreases B) increases C) remains the same  
7. As the concentration of reacting particles increases, the rate of reaction generally  
A) increased solvent contact B) increased solute solubility C) the equilibrium to shift to the left D) the equilibrium to shift to the right  
8. Given the reaction:

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## Topic 8 Kinetics And Equilibrium Answer Key

Equilibrium constant for the reaction,  $3A + 2B \rightleftharpoons C$  is.  $K =$

$$\left(\frac{C}{A^3 \times B^2}\right) 14.4$$

moles of A are mixed with 4 moles of B. At equilibrium for the reaction  $A + B \rightleftharpoons C + D$ , 2 moles of C and D are

formed. The equilibrium constant for the reaction will be: a)  $\left(\frac{1}{4}\right)$  b)

$\left(\frac{1}{2}\right)$  c) 1. d) 4. ANS:

## Chemical Equilibrium Important Questions And Answers

$$K_c = (2.0)(4.76 \times 10^{-31}) = 9.5 \times 10^{-31}.$$

The  $K_c$  values for each equilibrium in the sum are those appropriate to the ways in which they are written. Note that  $K_c$  for the first reaction in the sum is the inverse of the given value,  $1 / K_c$ , because it is being used in the reverse direction.

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